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Compound 1: 21.8 g C and 58.2 g O (80.0 - 21.8 = mass O) Compound 2: 34.3 g C and 45.7 g O (80.0 - 34.3 = mass O) The mass of carbon that combines with 1.0 g of oxygen is: Compound 1: Compound 2: 21.8 g C 58.2 g O 34.3 g C 45.7 g O = 0.375 g C/g O = 0.751 g C/g O The ratio of the masses of carbon that combine with 1 g of oxygen is  $0.751 = 0 \dots$

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8.71 Bonds broken: 1 C-N (305 kJ/mol) Bonds formed: 1 C-C (347 kJ/mol)  $\Delta H = \sum D_{\text{broken}} - \sum D_{\text{formed}}$ ,  $\Delta H = 305 - 347 = -42$  kJ Note: sometimes some of the bonds remain the same between reactants and products. To save time, only break and form bonds that are involved in the reaction.  
8.73 Bonds broken: Bonds formed:

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